

Name:

Period:

Seat#:

1) When calculating molarity, the volume needs to have what unit?	2) The maximum amount of solute dissolved is called _____.	3) Less than the maximum amount of solute dissolved is called _____.
4) More than the maximum amount of solute dissolved is called _____.	5) The solubility of solids goes _____ as the temperature is increased.	6) The solubility of gases goes _____ as the temperature is increased.
7) If you're trying to make a diluted solution, you use the equation:	8) When making a diluted solution the water added to the new solution is found by subtracting which two numbers?	9) Factors that affect rate are:
10) Factors that affect equilibrium position:	11) Only _____ changes the equilibrium constant (K_{eq})	12) What is average rate?

Dougherty Valley HS Chemistry
Things to Remember for Exam #2
Spring Test #2 – Solutions, Kinetics Equilibrium

<p>13) What is a rate expression? What is it used for?</p>	<p>14) When you want the rate of one substance but you only have the rate for another substance, you can use the _____ to solve for the missing rate. <i>Practice q: solve rate of h_2 in terms of n_2</i></p>	<p>15) The rate law only includes the concentrations of the _____.</p>
<p>16) The equilibrium expression is _____ divided by _____</p>	<p>17) The rate law exponents are called _____. Are they from the balanced equation coefficients or found experimentally?</p>	<p>18) Are the exponents in an equilibrium expression from the balanced equation coefficients or found experimentally?</p>
<p>19) Solids and liquids do or do not affect equilibrium?</p>	<p>20) A large value for k indicates that the _____ side is favored and a small value for k indicates the _____ side is favored.</p>	<p>21) $K'_{eq} = \text{????}$</p>
<p>22) If q is bigger than k, than the reaction will shift to the _____.</p>	<p>23) If q is smaller than k, than the reaction will shift to the _____.</p>	<p>24) I can use the 5% rule when:</p>
<p>25)</p>		